

## Electronic spectral studies and thermodynamic parameters of 5-Fluoro -1, 3- Dimethyl Uracil

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**Abstract:** The ultraviolet absorption spectra of this molecules have been recorded in various solvents. Thermodynamic parameters such as total energy, zero-point energy, rotational constants dipole moment, thermal energy, specific heat and entropy of 5-Fluoro-1, 3-Dimethyl Uracil are calculated. The plots of thermodynamic parameters have been given in following paper.

**Keywords:** Ultraviolet absorption, thermodynamic parameters, 5-1,3FDMU etc.

### I. Introduction

Uracil molecules are widely distributed in living organisms and find applications in biochemistry and clinical chemistry [1-3]. Recent spectroscopic studies of uracil and its derivatives have been motivated because of their biological and pharmaceutical importance. Uracil and its derivatives serves as allosteric regular and coenzymes for reactions in the human body and in plants. Some of its derivatives used as an anticancer drug (These drug inhibits RNA replication enzymes ,thereby diminating RNA synthesis and stopping the growth of cancerous cell ).Some of them are the basic constituents of DNA and RNA and play an important role in constitution properties of nucleic acids[4-5]. Recently computational methods based on density functional theory are becoming widely used. In the past two decades, ab initio quantum mechanical calculations of molecular structure and force field have contributed much to understanding of the vibrational spectra of organic compounds [6-9].

### II. Experimental Details:

The thermodynamic parameters viz enthalpy function  $[H^0-E_0^0]/T$ , heat capacity  $C_p^0$ , free energy  $[F^0-E_0^0]/T$  and entropy  $S^0$  at different temperatures are calculating from rigid rotator-harmonic oscillator approximation. The absorption spectrum of the compound was recorded in the range 200-400 nm with the Perkin Elmer Lambda 25 UV/VIS Spectrometer. were carried out at University scientific and instrumentation centre (USIC), Delhi University .

### III. Result And Discussion

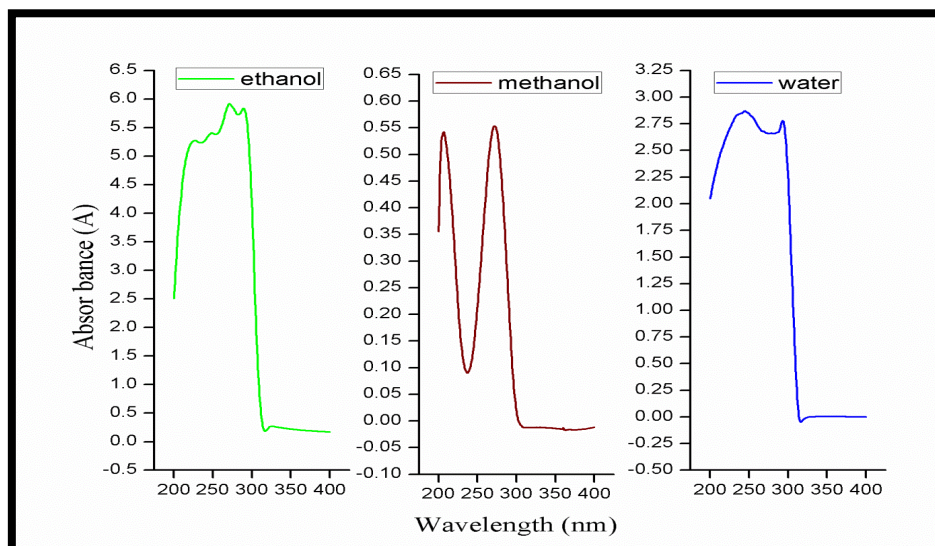
#### 3.1 Uv-Vis Absorption Spectra

In the visible and u.v. region the wavelength intensity and general appearance of the electronic bands, along with the effect of environmental conditions such as solvent and pH of the solutions are the important features for the characterization of compounds. The absorption spectra of pyrimidine bases show three or four absorption bands between 180-300nm. In addition to the long wavelength  $n-\pi^*$  transition, these bands appear in many pyrimidine bases and the presumably  $\pi-\pi^*$  transitions. The correlations indicate that the spectrum of each base is simply related to the spectrum of pyrimidine whose electronic states are derived from those of benzene. Thus benzene serves as the parent absorbing chromophore for all the pyrimidine bases[ 10-11]. Hence in view of the above discussion, the investigation has been undertaken to study the spectrum of 5-Fluoro-1,3-Dimethyl Uracil. The electronic spectra of this compound have been recorded in various polar solvents and their effect on electronic spectra have been discussed. In the present study, the UV spectra of 5-Fluoro-1, 3-Dimethyl Uracil were recorded in different solvents viz ethanol, methanol and water. The absorption spectrum of the molecule 5-Fluoro-1, 3-Dimethyl Uracil in the range 200-400 nm are given in Fig. 1.1. The observed bands of these molecules in ethanol, methanol and water are shown in Table 1.1. West [12] has suggested that changing of the solvent , effect the positions of bands in the electronic spectra which may be used to identify the bands as  $n-\pi^*$ ,  $\pi-\pi^*$  or  $n-\sigma^*$  . Generally the solvent used in recording these spectrum are non polar one (like hexane) or polar solvents (like ethanol, methanol and water). The band will be called blue shifted or red shifted according as the band shifts towards shorter wavelength or longer wavelength. On changing the solvent from hexane to water, the  $n-\pi^*$  transitions are attributed to blue shift bands, while  $\pi-\pi^*$  and  $n-\sigma^*$  transitions are attributed to red shift band . During the present investigation, the  $n-\pi^*$

are blue shifted in the title compound, with increasing polarity of the solvents i.e. (ethanol→ methanol→ water) as shown in the table 1.1 and in Fig. 1.1.

**Table 1.1 Solvent Effect on Electronic Transitions of 5 -Fluoro-1,3-Dimethyl Uracil**  
(All values are in nm)

Solvent	Refractive Index	Dielectric constant	$n - \pi^*$	$\pi - \pi^*$	$n - \sigma^*$
Ethanol	1.3773	25.0	271	262	226
Methanol	1.3362	32.0	272	-	206.5
Water	1.3380	80.5	278.5	240.5	-



**Fig. 1.1: Near ultraviolet absorption spectra of 5-fluoro-1,3-dimethyl uracil in ethanol, methanol and water**

### 3.2. Thermodynamic Parameters

The partition function and the statistical theory of enthalpy function  $[H^0 - E_0^0]/T$ , heat capacity  $C_p^0$ , free energy  $[F^0 - E_0^0]/T$  and entropy  $S^0$  have been discussed before optimization. To calculate the thermodynamic parameters one requires the principal moments of inertia, molecular weight, temperature and the vibrational fundamentals. The thermodynamic functions of 5-Fluoro-1,3-Dimethyl Uracil have been calculated with the help of computer program at different temperatures between 200-1500 K using the experimental fundamental frequencies assuming rigid rotor-harmonic oscillator approximation [13]. The thermodynamic parameters viz enthalpy function  $[H^0 - E_0^0]/T$ , heat capacity  $C_p^0$ , free energy  $[F^0 - E_0^0]/T$  and entropy  $S^0$  at different temperatures are given in Tables 1.2. The plots of thermodynamic parameters have been given in Fig. 1.2 and it was found that the thermodynamic functions rise more rapidly in the low temperature range and less rapidly in the high temperature range. At low temperatures only translational parts of motion gives its contribution but as temperature increases rotational and vibrational motions are excited [14]. Beyond a certain value of temperature (nearly at 1000 K) there is no further increase in molecular motion and therefore heat capacity becomes almost constant.

**Table 1.2: Thermodynamic Functions of 5-Fluoro-1,3-Dimethyl Uracil (in Cal. /Mol-K).**

Temperature K	Enthalpy $(H^0 - E_0^0)/T$	Heat Capacity $C_p^0$	Free energy $-(F^0 - E_0^0)/T$	Entropy $S^0$
200	9.61	14.20	103.85	113.46
300	12.29	21.13	108.24	120.53
400	15.36	27.99	112.19	127.55
500	18.56	34.62	115.96	134.51
600	21.76	40.92	119.63	141.39
700	24.93	46.78	123.22	148.14
800	27.99	52.12	126.75	154.74
900	30.95	56.93	130.22	161.17
1000	33.76	61.22	133.63	167.39
1100	36.44	65.02	136.97	173.41
1200	38.96	68.38	140.25	179.21
1300	41.34	71.34	143.46	184.80
1400	43.58	73.95	146.61	190.19

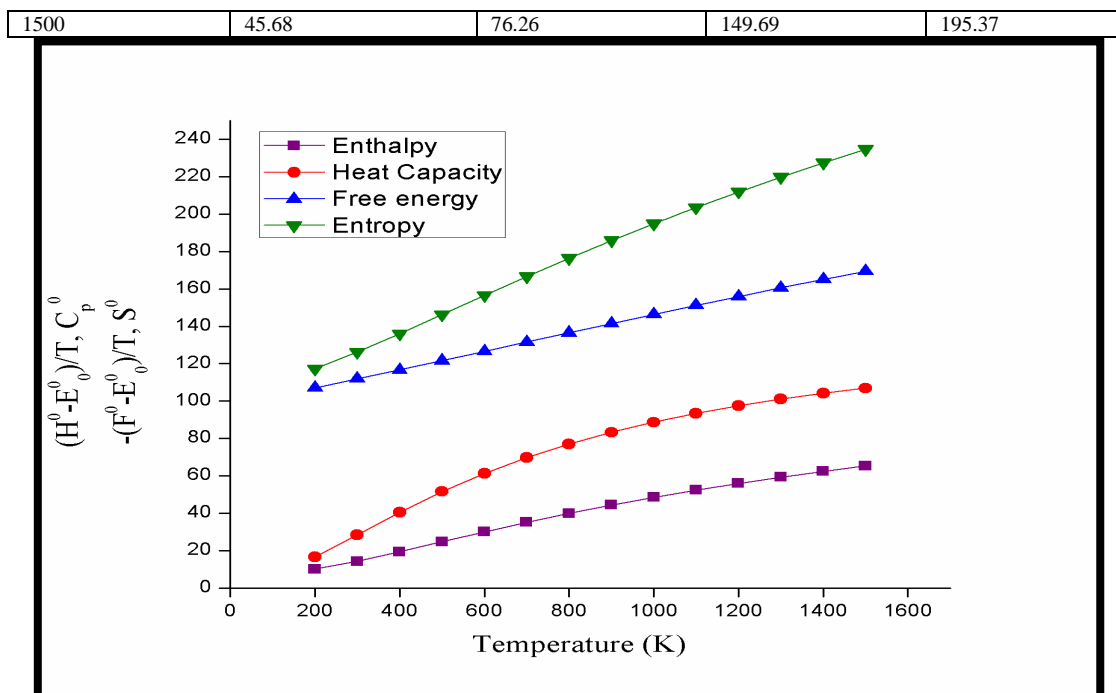


Fig. 1.2: Temperature V/s enthalpy, heat capacity, free energy and entropy of 5-fluoro 1, 3-dimethyl uracil

#### IV. Conclusion

The absorption spectrum, thermodynamic parameters viz enthalpy function  $[H^0-E_0^0]/T$ , heat capacity  $C_p^0$ , free energy  $[F^0-E_0^0]/T$  and entropy  $S^0$  at different temperatures are in good agreement with standard results.

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